

Acta Crystallographica Section E

Structure Reports

Online

ISSN 1600-5368

5-Chloromethyl-3-(3,5-dimethoxyphenyl)-1,2,4-oxadiazole

Pin-liang Wang, Hai-su Zeng, Hai-bo Wang* and Wen-yuan Wu

College of Science, Nanjing University of Technology, Xinmofan Road No. 5 Nanjing, Nanjing 210009, People's Republic of China
Correspondence e-mail: wanghaibo@njut.edu.cn

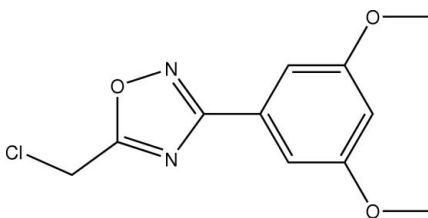
Received 19 September 2007; accepted 1 October 2007

Key indicators: single-crystal X-ray study; $T = 293$ K; mean $\sigma(\text{C}-\text{C}) = 0.005$ Å; R factor = 0.054; wR factor = 0.173; data-to-parameter ratio = 14.6.

In the title compound, $\text{C}_{11}\text{H}_{11}\text{ClN}_2\text{O}_3$, the substituted phenyl ring and the oxadiazole system are essentially coplanar. Close intramolecular $\text{C}-\text{H}\cdots\text{N}$ interactions are observed.

Related literature

For related literature, see: Nicolaides *et al.* (1998); Romero (2001); Terashita *et al.* (2002).



Experimental

Crystal data

$\text{C}_{11}\text{H}_{11}\text{ClN}_2\text{O}_3$
 $M_r = 254.67$
Triclinic, $P\bar{1}$
 $a = 7.861$ (2) Å
 $b = 8.502$ (2) Å

$c = 9.537$ (2) Å
 $\alpha = 80.72$ (3)°
 $\beta = 73.90$ (3)°
 $\gamma = 70.09$ (3)°
 $V = 574.2$ (2) Å³

$Z = 2$
Mo $K\alpha$ radiation
 $\mu = 0.33$ mm⁻¹

$T = 293$ (2) K
 $0.40 \times 0.20 \times 0.20$ mm

Data collection

Enraf-Nonius CAD-4 diffractometer
Absorption correction: ψ scan (North *et al.*, 1968)
 $T_{\text{min}} = 0.879$, $T_{\text{max}} = 0.937$
2422 measured reflections

2249 independent reflections
1589 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.031$
3 standard reflections every 200 reflections
intensity decay: none

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.054$
 $wR(F^2) = 0.173$
 $S = 1.08$
2249 reflections

154 parameters
H-atom parameters constrained
 $\Delta\rho_{\text{max}} = 0.31$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.34$ e Å⁻³

Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{C9}-\text{H9A}\cdots\text{N2}$	0.93	2.60	2.925 (5)	101
$\text{C5}-\text{H5A}\cdots\text{N1}$	0.93	2.54	2.838 (5)	99

Data collection: *CAD-4 Software* (Enraf-Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *SHELXTL* (Siemens, 1996); software used to prepare material for publication: *SHELXL97*.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: IM2037).

References

- Enraf-Nonius (1989). *CAD-4 Software*. Version 5.0. Enraf-Nonius, Delft, The Netherlands.
- Harms, K. & Wocadlo, S. (1995). *XCAD4*. University of Marburg, Germany.
- Nicolaides, D. N., Fylaktakidou, K. C., Litinas, K. E. & Hadjipavlou-Litina, D. (1998). *Eur. J. Med. Chem.* **33**, 715–724.
- North, A. C. T., Phillips, D. C. & Mathews, F. S. (1968). *Acta Cryst.* **A24**, 351–359.
- Romero, J. R. (2001). *Expert Opin. Investig. Drugs*, **10**, 369–379.
- Sheldrick, G. M. (1997). *SHELXL97* and *SHELXS97*. University of Göttingen, Germany.
- Siemens (1996). *SHELXTL*. Version 5.06. Siemens Analytical X-ray Instruments Inc., Madison, Wisconsin, USA.
- Terashita, Z., Naruo, K. & Morimoto, S. (2002). *PCT Int. Appl.* WO 0 260 439.

supplementary materials

Acta Cryst. (2007). E63, o4234 [doi:10.1107/S1600536807048064]

5-Chloromethyl-3-(3,5-dimethoxyphenyl)-1,2,4-oxadiazole

P. Wang, H. Zeng, H. Wang and W. Wu

Comment

1,2,4-Oxadiazoles are an important class of five-membered heterocyclic compounds. Some derivatives of 1,2,4-oxadiazoles showed intrinsic analgesic (Terashita *et al.*, 2002), antiinflammatory (Nicolaides *et al.*, 1998) and antipicornaviral (Romero, 2001) properties. We report here the crystal structure of the title compound, (I). The molecular structure of (I) is shown in Fig. 1, where the dashed lines indicate close C—H...N contacts which most probably are caused by the coplanar arrangement of the phenyl group and the oxadiazole system (Table 1). The other bond lengths and angles are of expected values (*cf.* Supplementary Material).

Experimental

A solution of chloroacetylchloride (16 mmol) in toluene (10 ml) was added dropwise to a solution of 3,5-(Dimethoxy)benzamidoxime (14 mmol) in toluene (60 ml). The resulting mixture was refluxed for 6 h, then concentrated under reduced pressure to afford crude compound(I). Pure compound (I) was obtained by recrystallization from a mixture of ethyl acetate (15 ml) and light petrolum (b.p. 60–90°C, 7.5 ml). Crystals of (I) suitable for X-ray diffraction were obtained by slow evaporation of an ethanol solution. ¹H NMR (CDCl₃, δ, p.p.m.): 6.55–6.56 (m, 2H), 6.23–6.25 (m, 1H), 4.64 (s, 2H), 3.77 (s, 6H).

Refinement

All H atoms bonded to the C atoms were placed geometrically at distances of 0.93–0.97 Å and included in the refinement using an riding motion approximation with $U_{\text{iso}}(\text{H}) = 1.2$ or $1.5U_{\text{eq}}$ of the carrier atom.

Figures

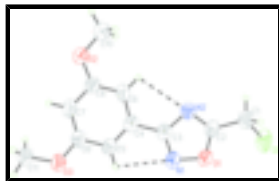


Fig. 1. A view of the molecular structure of (I), showing displacement ellipsoids at the 30% probability level. Dashed lines indicate C—H...N interactions.

5-Chloromethyl-3-(3,5-dimethoxyphenyl)-1,2,4-oxadiazole

Crystal data

C₁₁H₁₁ClN₂O₃

$M_r = 254.67$

Triclinic, $P\bar{1}$

$Z = 2$

$F_{000} = 264$

$D_x = 1.473 \text{ Mg m}^{-3}$

supplementary materials

Hall symbol: -P 1

$a = 7.861 (2) \text{ \AA}$

$b = 8.502 (2) \text{ \AA}$

$c = 9.537 (2) \text{ \AA}$

$\alpha = 80.72 (3)^\circ$

$\beta = 73.90 (3)^\circ$

$\gamma = 70.09 (3)^\circ$

$V = 574.2 (2) \text{ \AA}^3$

Mo $K\alpha$ radiation

$\lambda = 0.71073 \text{ \AA}$

Cell parameters from 25 reflections

$\theta = 9\text{--}13^\circ$

$\mu = 0.33 \text{ mm}^{-1}$

$T = 293 (2) \text{ K}$

Block, colourless

$0.40 \times 0.20 \times 0.20 \text{ mm}$

Data collection

Enraf–Nonius CAD-4
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 293(2) \text{ K}$

$\omega/2\theta$ scans

Absorption correction: ψ scan
(North *et al.*, 1968)

$T_{\min} = 0.879$, $T_{\max} = 0.937$

2422 measured reflections

2249 independent reflections

1589 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.031$

$\theta_{\max} = 26.0^\circ$

$\theta_{\min} = 2.2^\circ$

$h = -9 \rightarrow 9$

$k = -10 \rightarrow 10$

$l = 0 \rightarrow 11$

3 standard reflections

every 200 reflections

intensity decay: none

Refinement

Refinement on F^2

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.054$

$wR(F^2) = 0.173$

$S = 1.08$

2249 reflections

154 parameters

Primary atom site location: structure-invariant direct
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring
sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.05P)^2 + P]$$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.31 \text{ e \AA}^{-3}$

$\Delta\rho_{\min} = -0.34 \text{ e \AA}^{-3}$

Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > 2\sigma(F^2)$ is used only for calculat-

ing R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
Cl	0.3102 (2)	0.49956 (16)	0.15386 (12)	0.0816 (5)
O1	0.3892 (4)	0.1611 (3)	0.3056 (3)	0.0548 (7)
O2	-0.1673 (4)	-0.1007 (4)	0.9489 (3)	0.0684 (9)
O3	0.3152 (4)	-0.5222 (3)	0.6710 (3)	0.0578 (7)
N1	0.3733 (5)	0.0065 (4)	0.3778 (4)	0.0538 (8)
N2	0.1641 (4)	0.2229 (4)	0.5042 (3)	0.0451 (7)
C1	0.2542 (6)	0.4563 (5)	0.3446 (4)	0.0577 (10)
H1B	0.3402	0.4823	0.3861	0.069*
H1C	0.1297	0.5283	0.3852	0.069*
C2	0.2623 (5)	0.2789 (5)	0.3874 (4)	0.0448 (8)
C3	0.2402 (5)	0.0511 (4)	0.4938 (4)	0.0414 (8)
C4	0.1792 (5)	-0.0743 (5)	0.6045 (4)	0.0425 (8)
C5	0.2710 (5)	-0.2428 (4)	0.5881 (4)	0.0444 (8)
H5A	0.3678	-0.2777	0.5065	0.053*
C6	0.2172 (5)	-0.3588 (4)	0.6944 (4)	0.0446 (8)
C7	0.0694 (5)	-0.3066 (5)	0.8134 (4)	0.0492 (9)
H7A	0.0335	-0.3858	0.8841	0.059*
C8	-0.0273 (5)	-0.1346 (5)	0.8283 (4)	0.0478 (9)
C9	0.0298 (5)	-0.0180 (5)	0.7237 (4)	0.0469 (9)
H9A	-0.0304	0.0961	0.7325	0.056*
C10	-0.2766 (6)	0.0704 (5)	0.9669 (5)	0.0689 (12)
H10A	-0.3717	0.0780	1.0563	0.103*
H10B	-0.3335	0.1144	0.8859	0.103*
H10C	-0.1981	0.1342	0.9707	0.103*
C11	0.2735 (6)	-0.6468 (5)	0.7836 (4)	0.0576 (10)
H11A	0.3507	-0.7563	0.7535	0.086*
H11B	0.1447	-0.6389	0.8003	0.086*
H11C	0.2970	-0.6286	0.8722	0.086*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
Cl	0.1243 (11)	0.0765 (8)	0.0512 (6)	-0.0560 (8)	-0.0131 (6)	0.0155 (5)
O1	0.0591 (16)	0.0518 (16)	0.0486 (15)	-0.0255 (13)	0.0038 (12)	-0.0016 (12)
O2	0.083 (2)	0.0516 (17)	0.0531 (17)	-0.0204 (15)	0.0089 (15)	-0.0010 (13)
O3	0.0741 (19)	0.0412 (15)	0.0502 (15)	-0.0192 (13)	-0.0034 (13)	0.0006 (12)
N1	0.060 (2)	0.0454 (18)	0.0511 (19)	-0.0211 (15)	-0.0039 (16)	0.0029 (14)
N2	0.0504 (18)	0.0469 (18)	0.0389 (16)	-0.0189 (14)	-0.0094 (13)	-0.0001 (13)
C1	0.070 (3)	0.051 (2)	0.051 (2)	-0.024 (2)	-0.012 (2)	0.0041 (18)
C2	0.047 (2)	0.051 (2)	0.0414 (19)	-0.0225 (17)	-0.0130 (16)	0.0006 (16)
C3	0.0431 (19)	0.045 (2)	0.0422 (19)	-0.0196 (15)	-0.0157 (15)	0.0013 (15)
C4	0.0450 (19)	0.051 (2)	0.0402 (18)	-0.0238 (16)	-0.0155 (15)	0.0024 (15)

supplementary materials

C5	0.047 (2)	0.050 (2)	0.0375 (18)	-0.0187 (16)	-0.0077 (15)	-0.0029 (15)
C6	0.055 (2)	0.042 (2)	0.0414 (19)	-0.0190 (16)	-0.0166 (16)	0.0010 (15)
C7	0.063 (2)	0.046 (2)	0.042 (2)	-0.0247 (18)	-0.0132 (18)	0.0063 (16)
C8	0.056 (2)	0.050 (2)	0.0375 (19)	-0.0200 (17)	-0.0089 (16)	-0.0004 (15)
C9	0.058 (2)	0.0422 (19)	0.046 (2)	-0.0206 (17)	-0.0179 (17)	0.0043 (15)
C10	0.076 (3)	0.058 (3)	0.061 (3)	-0.019 (2)	-0.002 (2)	-0.006 (2)
C11	0.072 (3)	0.041 (2)	0.052 (2)	-0.0175 (19)	-0.006 (2)	0.0038 (17)

Geometric parameters (Å, °)

C1—C1	1.757 (4)	C4—C9	1.397 (5)
O1—C2	1.322 (4)	C5—C6	1.380 (5)
O1—N1	1.411 (4)	C5—H5A	0.9300
O2—C8	1.343 (4)	C6—C7	1.384 (5)
O2—C10	1.426 (5)	C7—C8	1.410 (5)
O3—C6	1.360 (4)	C7—H7A	0.9300
O3—C11	1.434 (4)	C8—C9	1.384 (5)
N1—C3	1.298 (5)	C9—H9A	0.9300
N2—C2	1.290 (4)	C10—H10A	0.9600
N2—C3	1.386 (4)	C10—H10B	0.9600
C1—C2	1.482 (5)	C10—H10C	0.9600
C1—H1B	0.9700	C11—H11A	0.9600
C1—H1C	0.9700	C11—H11B	0.9600
C3—C4	1.478 (5)	C11—H11C	0.9600
C4—C5	1.380 (5)		
C2—O1—N1	106.2 (3)	O3—C6—C7	124.1 (3)
C8—O2—C10	117.5 (3)	C5—C6—C7	120.3 (3)
C6—O3—C11	117.6 (3)	C6—C7—C8	120.6 (3)
C3—N1—O1	103.1 (3)	C6—C7—H7A	119.7
C2—N2—C3	101.7 (3)	C8—C7—H7A	119.7
C2—C1—C1	112.7 (3)	O2—C8—C9	126.2 (4)
C2—C1—H1B	109.1	O2—C8—C7	114.7 (3)
C1—C1—H1B	109.1	C9—C8—C7	119.1 (3)
C2—C1—H1C	109.1	C8—C9—C4	119.1 (3)
C1—C1—H1C	109.1	C8—C9—H9A	120.5
H1B—C1—H1C	107.8	C4—C9—H9A	120.5
N2—C2—O1	114.5 (3)	O2—C10—H10A	109.5
N2—C2—C1	126.6 (3)	O2—C10—H10B	109.5
O1—C2—C1	118.8 (3)	H10A—C10—H10B	109.5
N1—C3—N2	114.5 (3)	O2—C10—H10C	109.5
N1—C3—C4	121.6 (3)	H10A—C10—H10C	109.5
N2—C3—C4	123.9 (3)	H10B—C10—H10C	109.5
C5—C4—C9	121.8 (3)	O3—C11—H11A	109.5
C5—C4—C3	119.4 (3)	O3—C11—H11B	109.5
C9—C4—C3	118.7 (3)	H11A—C11—H11B	109.5
C6—C5—C4	119.1 (3)	O3—C11—H11C	109.5
C6—C5—H5A	120.5	H11A—C11—H11C	109.5
C4—C5—H5A	120.5	H11B—C11—H11C	109.5
O3—C6—C5	115.6 (3)		

C2—O1—N1—C3	-0.4 (4)	C3—C4—C5—C6	177.8 (3)
C3—N2—C2—O1	0.6 (4)	C11—O3—C6—C5	175.6 (3)
C3—N2—C2—C1	-174.9 (4)	C11—O3—C6—C7	-6.0 (5)
N1—O1—C2—N2	-0.1 (4)	C4—C5—C6—O3	-179.5 (3)
N1—O1—C2—C1	175.7 (3)	C4—C5—C6—C7	2.0 (5)
Cl—C1—C2—N2	-150.1 (3)	O3—C6—C7—C8	-178.9 (3)
Cl—C1—C2—O1	34.6 (5)	C5—C6—C7—C8	-0.5 (6)
O1—N1—C3—N2	0.8 (4)	C10—O2—C8—C9	4.6 (6)
O1—N1—C3—C4	-178.7 (3)	C10—O2—C8—C7	-177.0 (4)
C2—N2—C3—N1	-0.9 (4)	C6—C7—C8—O2	-179.7 (4)
C2—N2—C3—C4	178.6 (3)	C6—C7—C8—C9	-1.2 (6)
N1—C3—C4—C5	3.0 (5)	O2—C8—C9—C4	179.7 (4)
N2—C3—C4—C5	-176.4 (3)	C7—C8—C9—C4	1.3 (5)
N1—C3—C4—C9	-177.3 (4)	C5—C4—C9—C8	0.2 (5)
N2—C3—C4—C9	3.3 (5)	C3—C4—C9—C8	-179.5 (3)
C9—C4—C5—C6	-1.9 (5)		

Hydrogen-bond geometry (Å, °)

<i>D—H...A</i>	<i>D—H</i>	<i>H...A</i>	<i>D...A</i>	<i>D—H...A</i>
C9—H9A...N2	0.93	2.60	2.925 (5)	101
C5—H5A...N1	0.93	2.54	2.838 (5)	99

Fig. 1

